



1 Intermolecular Forces

Polar Interactions

1 Hydrogen Bonding

Van Der Waal Interactions

1 The Amide Bond

□ Bonding Interactions in Tertiary Structure
 □ lonic

Covalent

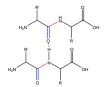
1 Examples of Proteins



Polar Interactions

Stabilizing force arising from the attraction between positive and negative portions of protein chain. The more polar the portions, the stronger the attractive force.

Special type of polar interaction: The hydrogen bond strong polar interaction between a somewhat electron deficient hydrogen and an electronegative (electron rich) oxygen or nitrogen.



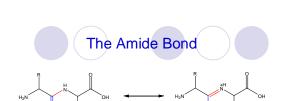




Stabilizing force present between all molecules/portions of protein. The movement of electrons cause temporary dipoles to form, which induce other dipoles. The dipoles attract.



http://www.chem.unsw.edu.au/coursenotes/CHEM1/nonunipass/HainesIMF/dispersion.html



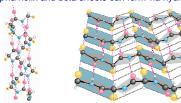
- Amide has resonance structure
- Electronics of amide can be seen in resonance structure



Consequences

Omega bond is either trans (180 degrees) or cis (0 degrees)

1 Alpha helix and beta sheets can form via hydrogen bonding



Bonding Interactions in Tertiary Structure

Lonic Bond

Bond formed when two atoms are held together by electrostatic forces.

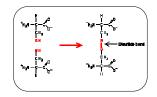
Bonds can form between positively and negatively charged side



Bonding Interactions in Tertiary Structure

Covalent Bond

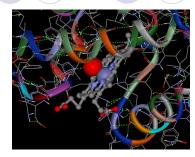
Bond formed when electrons are shared between atoms Example: disulfide bridge





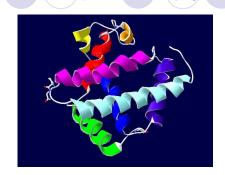
http://chemistry.umeche.maine.edu/CHY431/Proteins11.html

Heme: a cofactor



http://www.accelrys.com/pharma/targe

Protein Examples: Horse myoglobin



Ramachandran Plot

